

**Calibration of
atmospheric
hydrogen**

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Calibration of atmospheric hydrogen measurements

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Abstract

Interest in atmospheric hydrogen (H_2) has been growing in recent years with the prospect of H_2 being a potential alternative to fossil fuels as an energy carrier. This has intensified research for a quantitative understanding of the atmospheric hydrogen cycle and its total budget, including the expansion of the global atmospheric measurement network. However, inconsistencies in published observational data constitute a major limitation in exploring such data sets. The discrepancies can be mainly attributed to difficulties in the calibration of the measurements. In this study various factors that may interfere with accurate quantification of atmospheric H_2 were investigated including drifts of standard gases in high pressure cylinders. As an experimental basis a procedure to generate precise mixtures of H_2 within the atmospheric concentration range was established. Application of this method has enabled a thorough linearity characterization of the commonly used GC-HgO reduction detector. We discovered that the detector response was sensitive to the composition of the matrix gas. Addressing these systematic errors, an accurate calibration scale has been generated defined by thirteen standards with dry air mole fractions ranging from 139–1226 nmol mol⁻¹. The new scale has been accepted as the official World Meteorological Organisation's (WMO) Global Atmospheric Watch (GAW) H_2 mole fraction scale.

1 Introduction

Following some early measurements of atmospheric hydrogen (Schmidt, 1974) this trace gas has received little attention for many years. In the past decade, interest in the atmospheric H_2 budget has increased significantly as scenarios for a future hydrogen fuel economy have been developed (Prather, 2003). The main concern about rising levels of atmospheric hydrogen is an additional consumption of hydroxyl radicals which in turn would influence the lifetime of greenhouse gases such as methane. Growing H_2 mixing ratios would also increase stratospheric water vapour concentrations with

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(Masarie et al, 2001; Bonasoni et al., 1997), which are likely to have contributed to inconsistencies in various concentration time series.

All published atmospheric H₂ measurements have been carried out using gas chromatography with a mercuric oxide reduction detector. Only very recently, alternative technologies suited for routine operation have been described (Novelli et al., 2009; Necki et al., 2009). The HgO reduction detector has a non-linear response, which demands a careful characterization using multiple calibration points. In turn, this makes the quantification sensitive to the applied response function and the concentration range of the calibration standards.

In view of the described difficulties associated with calibration of H₂ measurements a European research project for investigating the biogeochemistry of H₂ (EuroHydros) included activities to set up an accurate calibration scale for improving the consistency of the various data sets. Studies that were performed within this project included long-term background air monitoring (Grant et al., 2010; Yver et al., 2010), observations in urban environments (Vollmer et al., 2007; Aalto et al., 2009; Hammer et al. 2009; Yver et al., 2009) and soil uptake studies (Hammer and Levin, 2009; Lallo et al., 2009). Hydrogen mixing ratios detected in these experiments span from the detection limit to more than 1000 ppb. Here we describe the development of a new calibration scale ranging from 139 to 1226 nmol mol⁻¹ (parts-per-billion, ppb).

2 Experimental section

Three conditions are required for the accurate determination of hydrogen in air: a precise analytical method, stable reference standards and an accurate assignment of H₂ mixing ratios to these standards.

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2.1 Analytical method

Air samples were measured using a Reduction Gas Analyser (RGA) 3 (Trace Analytical, Menlo Park (CA), USA) that is based on gas chromatographic separation with subsequent HgO reduction and photometric UV detection of Hg (Schmidt and Seiler, 1970). The instrument is equipped with a pre-column of Unibeads 1S (1/8" × 0.8 m, 60/80 mesh) connected in series with an analytical column containing Molecular Sieve 5Å (1/8" × 0.8 m, 60/80 mesh) through a 2-position, 10-port injection valve (ET6C10UWE, VICI, Schenk, CH) and a 1 mL sample loop. The 10-port injection valve configuration allows the pre-column to be back flushed, thus preventing other reducible trace gases from reaching the analytical column and detector. Synthetic air (Westfalen AG, Muenster, Germany) is used as the carrier gas at a flow of 25 mL/min. The GC and detector ovens are kept constant at 106 °C and 272 °C, respectively. The sample is first flushed through the sample loop for 36 s at a flow of 70 mL/min. After the sample flow is stopped the sample loop pressure is allowed to relax to ambient pressure for 20 s before the injection valve (ET6C10UWE, VICI, Schenk, CH) is switched to start the analysis. The H₂ retention time is 49 s, CO elutes at 104 s. The flow direction in the pre-column is reversed at 78 s. Because of the characteristic baseline noise and tailing peak shape of the detector the performance of the peak integration algorithm is crucial for the overall performance. The analogue output from the RGA3 is digitized using an analogue-to-digital converter (35900E, Agilent Technologies, Santa Clara (CA), USA) and integrated using the Agilent Chemstation software (Version A.08.03). Quantification is based on the sample peak height normalized to the mean H₂ peak height of the bracketing analysis pair of the working reference gas.

Two external gas selection valves allow 6 standards and an additional 16 flasks to be analyzed in series, always alternating with a gas aliquot from a single working reference cylinder. For the period of the experiments the raw peak heights of this working reference were stable within ±3.3%. The HgO reaction bed had been in use for five years. Major response changes have been observed after power failures but

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no such incident had occurred since September 2007. Three of the standard ports are permanently occupied by quality control air standards at 500, 650 and 1200 ppb that are generally analyzed on a daily basis.

2.2 Air standard production

5 Measurements of atmospheric hydrogen are made relative to a set of standard gases in high pressure cylinders. These standards are natural air samples, filled at the Max-Planck-Institute for Biogeochemistry using an oil-free compressor (Rix Sweet Air 6A 5-L, Rix Industries, Benicia (CA), USA) with anhydrous magnesium perchlorate as the drying agent. To generate standards with a sub-ambient H₂ mixing ratio air was de-
10 pleted of H₂ by directing the compressor flow through a cartridge filled with a hydrogen scrubber (Sofnocat 423 (775 g) plus Sofnocat 514 (1000 g), Molecular Products, Thaxed, UK). Standards with higher concentrations were produced by admixture of varying amounts of commercial compressed air (Linde AG, Leipzig, Germany) containing a high H₂ mixing ratio. The H₂ mixing ratios of these standards have been
15 initially assigned relative to a single standard gas with 543 ppb H₂ provided by CSIRO, Division of Marine and Atmospheric Research, Aspendale, Australia in 2002. This enabled the link to the H₂ CSIRO94 calibration scale used by the AGAGE monitoring network (Francey et al., 1996). The preliminary characterization of the non-linearity of the HgO-reduction detector response was made using a set of 18 flask samples that
20 were prepared with variable proportions of a standard air containing about 800 ppb H₂ and 3000 ppb CH₄, respectively, and synthetic, H₂- and CH₄-free air. The H₂ mixing ratios ranged from 100 to 800 ppb, with 6 data points in the atmospheric range (400–600 ppb). The dilution factor for the individual samples was calculated from the ratio of the CH₄ concentration in the sample to the CH₄ concentration in the undiluted stan-
25 dard. Methane was analyzed by a GC flame ionisation detector (FID) which exhibits a good linearity in the specified concentration range. With a quadratic response function established for the HgO-reduction detector in this experiment the H₂ mixing ratios of the standard gases in high pressure cylinders were preliminarily assigned (see column

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“CSIRO94” in Table 1). As one approach to verify the stability of the standards, aliquots from five of these standards were filled into low pressure 5-L glass flasks equipped with one valve with PCTFE sealing. This amount of sample has enabled repeated analysis over five years, which allowed to monitor the relative stability of H₂ concentrations of the flask samples compared to the set of high-pressure cylinder standard gases.

2.3 Method description: set-up for generating new standard gas mixtures

With the procedure described in the section above a set of calibration standards was produced that ensured a consistent data set. However, the various calibration transfer steps as well as the unknown uncertainty of the CSIRO94 calibration scale made it difficult to assess its accuracy. In order to achieve assignments with better traceable accuracies, a method to precisely produce H₂ gas mixtures was set up and evaluated. The technique described in this section is based on dilution of a known volume of pure hydrogen with a known mass of hydrogen-free real air that has been prepared using the compressor system described in the previous section. The volume is defined by a stainless steel sample loop connected to a two position, 10 port valve (ET6C10UWM, VICI, Schenkon, CH). The volume determination is made as follows: the sample loop is quantitatively filled with degassed, high-purity water and weighed several times on a calibrated balance (AT261, Mettler-Toledo, Giessen, Germany) until weight loss due to evaporation is apparent. To calculate the volume the density of water at the measured temperature is used. The mass of the displaced air in the sample loop is also accounted for (corresponding to 0.11% of the liquid water mass). For assessing the reproducibility of this measurement the filling and weighing procedure was performed 10 times in a row and the whole procedure was repeated on three different days. When the sample loop is connected to the Valco valve an additional volume of 2.46 μL has to be accounted for, which results from the channel in the valve rotor and from the boreholes of the employed valve (Kurrmann, 2005). Most experiments were performed using a sample loop of 344.6 μL volume. Before starting the dilution of hydrogen the amount of diluent air needed for the targeted mixing ratio is calculated for the given

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amount of H₂. The number of moles of H₂ is determined according to the gas law (taking into account non ideal behaviour of the pure gas):

$$n(\text{H}_2) = \frac{p \cdot V}{R \cdot T} \cdot \left(1 + \frac{B}{V_{\text{mol}}} \right) \quad (1)$$

with p = atmospheric pressure [Pa], V = volume of sample loop + internal volume of the valve [L]; R = ideal gas constant = 8314.51 (Pa*L)/(K*mol); T = temperature [K]; B = virial coefficient for H₂ = 0.015 L/mol (Kehiaian, 1997); V_{mol} = molar volume = $R \cdot T / p$.

The corresponding mass of diluent air (m_{air}) needed for a targeted mixing ratio then corresponds to:

$$m_{\text{air}} = M_{\text{air}}^* n_{\text{H}_2} / \text{mixing ratio} \quad (2)$$

with M_{air} = molar mass of air = 28.965 g/mol (Tohjima et al., 2005; Wieser et al., 2009)

Initially, sample loop and valve are stored for a few hours in a thermo insulating foam box to ensure a homogeneous temperature of sample loop, valve and a PT100 thermocouple (GTF 175, Greisinger electronic, Regenstauf, Germany). The two valve ports adjacent to the ports that are occupied by the sample loop are coupled via stainless steel capillaries (1 m each, 1/16" OD, "ID; Valco, Schenk, CH) with a hydrogen generator (Parker-Balston 75-32, Haverhill (MA), USA) and a vent, respectively. The procedure consists of the following steps:

H₂ is filled in the sample loop by flushing it for 15 s with a total volume of >100 mL of pure hydrogen. The flow is then stopped, and following a pressure equilibration time of 20 s, the valve is switched to isolate the sample loop contents. The temperature of the sample loop and the ambient pressure measured using a high precision barometer (DPI142, Druck GE Sensing, Leicester, UK) are recorded. Subsequently, the interior of the valve is flushed with several liters of N₂ to ensure that no H₂ other than the amount captured inside the sample loop remains inside the valve. The two capillaries that were in contact with pure H₂ are detached. Next, the valve is installed in a panel

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required mass for the targeted mixing ratio is reached. After removal of the transfer tubing the cylinder mass is weighed to determine the exact amount of added diluent gas.

Finally, the gas mixtures are analysed on the GC-HgO detector repeatedly. After analysis some of the mixtures are partly blown off through a 1/16" capillary, their mass is weighed and a new mixture is produced by dilution with additional hydrogen-free air. Analysis of non-diluted mixtures after the pressure release did not reveal any change of the H₂ mixing ratio due to a fractionation during the blowing-off process.

A series of 53 hydrogen-in-air mixtures in the range of 12–1252 ppb was prepared by the above described mixing technique during September–December 2008 and thoroughly analysed to characterize the response curve of the detector. A list of the individual mixtures and their analysis results is given in Table 2. To verify the stability of the detector response during this period several rounds of analysis of the 13 standard gases listed in Table 1 have been conducted in addition to the daily analysis of the quality control standard gases.

3 Results and discussion

3.1 Reference gas stability

Various container types have been tested for their suitability of storing H₂ standard gases. A precondition for this stability test is a fixed reference point that is known not to drift. As such reference the set of 13 calibration standards listed in Table 1 has been chosen. Support for the assumption that these standard gases are stable comes from the record of H₂ concentration measurements of aliquots from the 5-L glass flasks samples. No detectable change of H₂ was observed in these samples relative to the set of calibration standards over five years (Fig. 2). Additional evidence for the stability of the reference standard set is provided by results from an ongoing bi-monthly flask intercomparison exercise between MPI-BGC and CSIRO. The inter-laboratory offset of

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H₂ measurements has remained very stable throughout the same period of 2005–2009 (mean offset: CSIRO-MPI_{CSIRO94} = -0.3 ± 2.0 ppb, $n = 76$).

About 100 air standards in various types of high pressure cylinders were analysed repeatedly for periods of one to six years to judge the stability of their H₂ mixing ratios.

If an increase in the H₂ mixing ratio of more than 3 ppb is observed for air containing ambient levels of H₂ (400–650 ppb) this is assumed to reflect a true drift, while smaller differences may just reflect the longer term uncertainty of the measurements. This assessment criterion is based on the results of the routine quality control gas standard analysis. Over the period of April 2007 through August 2009 a standard deviation of ≤ 1.6 ppb was observed for the daily means of three standards with H₂ in the range from 500 to 680 ppb (Fig. 3). A larger scatter was observed for the quality control standard at 1200 ppb H₂ (std. dev. = 5 ppb). Hence, a H₂ drift in air standards exhibiting mixing ratios of 650–1200 ppb is assumed significant if the absolute change exceeds 1% of its H₂ mixing ratio.

Figure 4 summarizes the H₂ drift rates observed in sets of standards grouped by cylinder type relative to the reference set. High pressure cylinders made from steel or stainless steel generally kept stable H₂ levels while highly variable storage properties were observed in aluminium cylinders. For the aluminium cylinder type that is most commonly used for standards of greenhouse gases (N150; Luxfer, Riverside, CA, USA)(NOAA, 2010) only about 40% of the cylinders tested were found to keep stable H₂ mixing ratios; some cylinders showed very small H₂ drifts and others featured large H₂ increases. In addition, the hydrogen drift rates of standard gases in 5-L cylinders manufactured from different aluminium alloys have been compared. The most commonly used alloy AA6061 is an aluminium-magnesium-silicon blend (Aluminium Association, 2009). The cylinder manufactured from this material exhibited an increase of H₂ which is significantly higher and longer-lasting compared to that in cylinders of the alloys AA2001 (major alloying element Cu) and AA7060 (major alloying element Zn) (Fig. 5). In these latter containers the H₂ mixing ratio stabilized within a few months after an initial lower drift. These different properties reflect the influence of alloying

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conditions during the mixing experiments. Temperature variations in the weighing laboratory were always below 0.5°C , rendering the temperature influence on the density a minor effect. Other systematic errors may include the level of filling with water and the uncertainty of the internal volume of the valve (including fittings). The relative contribution of these errors to the total uncertainty would change with a variation of the sample loop size. However, experiments performed using alternative loops with different volumes ($245.8\ \mu\text{L}$, $379.9\ \mu\text{L}$, $1511\ \mu\text{L}$) yielded consistent results and did not reveal any systematic loop size dependence. The correction from the ideal gas law that is made by applying the second virial coefficient is very small (0.06%) and the uncertainties associated with this coefficient are considered to be negligible.

A blank determination of the diluent gases using the GC-HgO method did not provide a rigorous purity assessment, because of the non-sufficient detection limit of the analytical method (~ 10 ppb). In order to check for the completeness of the catalytic removal of hydrogen in the diluent gas, experiments for the dilution of H_2 were performed at different flow rates. Thus, the residence times of the diluent in the purifying cartridge were varied. An increase in H_2 concentrations that points to a non-quantitative removal of H_2 in the dilution air was observed at flow rates exceeding $5\ \text{L}/\text{min}$. Therefore, special care was taken that the flow rate of the gas transfer was always restricted to $2.5\ \text{L}/\text{min}$ at maximum. The absence of a significant blank in such dilution air has been verified using a new analytical system based on gas chromatography with a pulsed-discharge photo ionization detector that has been installed recently (GC-HePDD; VICI, Schenkon, CH) (Novelli et al., 2009). A chromatogram of diluent air with this instrumentation indicated a very small peak below the detection limit of <1 ppb. However, this test could only be done by hindsight as the system has been set up after the production series of standard mixtures had been completed.

The palladium diffusion cell that is utilized in the hydrogen generator is specified to reach a hydrogen purity of $>99.99999\%$. This excludes a significant contribution of any hydrogen impurity to the uncertainty of the method.

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In contrast, it is essential to exclude loss of H₂ or contamination with excess H₂ during any of the steps of the experiment. As pointed out earlier, a H₂ increase has been identified in air samples stored in aluminium cylinders. This renders this type of container unsuitable for long-term storage of reference gases. Therefore, the air mixtures generated with the mixing procedure were not kept as reference standards but served to assign values to standards in containers with superior storage properties. These latter cylinders have a larger mass exceeding the weighing range of the balance available which excluded their direct use for preparation of gravimetric mixtures. Therefore, the requirement for stable hydrogen mixing ratios in the 5-L aluminium cylinder applies only to the relatively short time of the experiment including the time for analysing the gas mixture. A storage test with an air sample in the cylinder used for the experiment was showing an initial H₂ growth rate of 0.04 ppb/d (depicted by triangles in Fig. 5). Within a maximum period of 5 days between generation of a mixture and its analysis this growth rate does not cause any detectable alteration of the composition.

There is also the potential for a loss of H₂ within the process. Significant differences in the permeation rates of hydrogen relative to oxygen or nitrogen have been described for various polymers including polyarylethers (Wang et al., 2002) and fluorinated polymers (Monson et al., 2009). One polymer part that is in contact with the H₂ during the preparation of the gas mixtures is the rotor of the two position valve. This may result in some H₂ diffusing into the valve rotor material while the sample loop is isolated between the filling and transfer to the evacuated cylinder, which takes about 6 min. To investigate this effect several mixtures were prepared with two different rotor polymer types and variable holding times. A decrease of the detected H₂ at longer waiting periods was observed for mixtures, which had been prepared using a valve equipped with a Valcon-E rotor (polyarylether/PTFE polymer composite; VICI, Schenk, CH). This finding suggests a loss of hydrogen by diffusion into this polymer. No such depletion was apparent for mixtures prepared with Valcon-M rotors (hydrocarbon polymer; VICI, Schenk, CH) for holding times up to two hours. Therefore, all standard mixtures discussed here were prepared with Valcon M-rotors.

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The residuals from the best regression fit of the GC-HgO-reduction measurement data of the set of 53 H₂-in-air mixtures (see Sect. 3.3) provide an experimental measure of the combined uncertainty contributed by the precision limits of the mixing procedure and the measurement error. In Table 2 the random components of the procedure's uncertainty (related to the precision limits of the pressure, temperature and mass determinations) are summed up for each mixture taking the square root of the sum of the individual error squares (Table 2, column: $u_{\text{mix random}}$). In the same table the analytical precision and the residuals of the regression fit are specified. The combined measurement and random mixing uncertainty is on average 0.12% for the mixtures with H₂ mixing ratios above 300 ppb which complies very well with the observed relative residuals. More relevant for the accuracy of the procedure are systematic errors associated with the measured parameters and other influencing factors. These do not average out with increasing sample number but can cause a systematic bias. Including the potential systematic errors identified in Table 3 results in an estimate of the relative uncertainty of 0.3% for the mixing procedure (see Table 2, last column).

3.3 HgO reduction detector response function

A series of 53 mixtures in the range of 12–1250 ppb was prepared by the above described mixing technique during September–December 2008 (Table 2). The relationship between signal response and actual mixing ratio is highly non-linear (Fig. 6). Systematic, concentration dependent residuals are apparent in 2nd and 3rd order polynomial regression functions (Fig. 7). In contrast, a combination of a quadratic and exponential function of the type:

$$[\text{H}_2] = ax^2 + bx + c*(1 - e^{-dx}) \quad (3)$$

a , b , c , d = regression parameters; x = peak height normalised to working standard.

Results in a good match between measured and calculated data points. This response function type is the result of the detection principle including an exponential

relationship from the Lambert-Beer law of the UV absorption detection of the mercury vapour as well as an influence from the chemical reaction of the HgO reduction.

In addition to the mixtures in air, some 50 mixtures were prepared with nitrogen, purified using a gas cleaning cartridge certified for high pressure and <1 ppb impurity levels (Aeronex 70KFI4R, Mykrolis, San Diego (CA), USA). Over the entire observed mixing ratio range the signal heights of H₂ mixing ratios in nitrogen were consistently 0.6% higher than in the corresponding H₂-in-air mixtures with the same H₂ content (Fig. 8). Additional mixtures of H₂ in argon, in oxygen, and mixtures with various fractions of nitrogen in air or oxygen in air show a clear anti-correlation of the H₂ signal response relative to the oxygen concentration (Fig. 9). This influence of oxygen on the detector sensitivity can be explained by an overlap of the chromatographic peaks of hydrogen and oxygen. Although oxygen does not directly generate any signal at the HgO reduction detector it can re-oxidize some of the Hg that is produced by the reduction of HgO through H₂. Thus, oxygen reduces the detector signal strength. This matrix sensitivity requires that quantification of H₂ in air samples be done by calibration of the detector using H₂-in-air mixtures as standard gases. Reference gases in nitrogen or any other gas would result in a bias.

Using the calibration function described above (excluding the lowest data point which was at the detection limit of the HgO-reduction detector) the H₂ mixing ratios of the suite of 13 standard gases of the reference set were reassigned (Table 1, column MPI-2009). Recent analysis of these gases using the GC-HePDD with a linear detector response resulted in absolute residuals below 1 ppb (see Table 1, last column). This finding provides independent support for the adequate description of the HgO-reduction detector response.

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4 Summary and conclusions

This study has investigated factors that are limiting the accuracy of atmospheric hydrogen measurements. Major limitations that have been identified include the availability of reference standards, the stability of the standard gas composition over time, the non-linearity of the prevalent detecting principle and its dependence on the composition of the standard gas matrix.

A great variability in the properties of different aluminium cylinders with respect to maintaining stable H₂ mixing ratios is observed ranging from no detectable drift to H₂ growth rates >20 ppb/yr. In contrast, steel cylinders are generally found to be suitable containers for H₂ reference gases.

A method for the production of precisely defined mixtures of H₂ in hydrogen free air or other gases has been set-up. The method is a straightforward alternative to the costly procedure of gravimetric preparations of calibration standards that has limited the rate of preparation of standards in many laboratories. It has been used to produce more than 100 individual H₂ standards in real air and other gases. Evaluation of the factors that influence the accuracy of the method results in an uncertainty estimate of 0.3% for the procedure. The analysis results of 52 air mixtures covering a concentration range of 50–1250 ppb were used to characterize the response function of the commonly used GC-HgO-reduction detector. Over this concentration range the response curve cannot be adequately described with a 3rd order polynomial function but is better characterized with a function with exponential and quadratic terms. Comparison of HgO-reduction analysis data of H₂ in gas mixtures with different oxygen contents revealed a dependence of the signal on the oxygen content of the matrix gas. This demonstrates the importance to calibrate the GC-HgO-reduction analyser with air mixtures.

The cylinder that was used for this gravimetric mixing needed to meet the mass limits for the balance but was not suitable for long-term storage of hydrogen in air mixtures. Therefore, these reference mixtures were discarded after thorough analysis. Their

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analysis results were used to assign hydrogen mixing ratios to 13 standard gases. In metrological terms these standards are no “primary” standards (Brown and Milton, 2007) as the GC analysis does not qualify as primary method. Yet, this set of standards fulfils the basic requirements for a calibration scale as it is stable in time, comparable and coherent with measurements using alternative technologies. This is verified by repeated analysis of these standards with the GC-HePDD technique yielding very small residuals applying a linear response fit.

The 13 standards constitute a new H₂ calibration scale that is identified as MPI-2009 H₂ scale. It has recently been accepted by the community of experts within the World Meteorological Organization (WMO) as WMO H₂ mole fraction scale and such serves as reference material for atmospheric observations within the Global Atmosphere Watch (GAW) program. This agreement is a first step towards obtaining a global, integrated database for H₂ monitoring data.

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Table 1. Set of reference standards.

Cylinder ID ^a	p [bar]	filling date	H ₂ [ppb]		σ^c [ppb]	PDD residuals ^d [ppb]
			CSIRO94 ^b	MPI-2009		
ESX-0803237	60	26.04.2007	134.6	139.4	0.4	-0.6
GM-26/470-3	22	13.02.2007	274.0	289.2	0.4	0.3
GM-24/635-22	22	07.02.2007	399.8	415.5	0.3	-0.1
ESX-0803222	60	22.08.2004	456.0	471.7	0.3	0.5
ESX-0803231	60	04.08.2004	492.4	507.8	0.3	0.2
Lin-B0892/92	200	10.06.2003	512.2	529.2	0.4	1.0
Lux-3502361	200	22.11.2004	562.4	581.1	0.7	0.1
Lin-B0873/92	200	22.02.2007	604.1	622.7	0.5	-0.7
ESX-0500421	60	10.12.2005	664.2	685.2	0.9	-0.5
GM-26/470-2	22	15.02.2007	731.4	754.0	1.1	-1.0
GM-24/635-2	22	07.02.2007	824.8	854.6	0.7	0.2
GM-26/470-1	22	09.02.2007	937.9	970.4	1.9	0.8
GM-24/635-30	22	05.04.2006	1185	1225.6	2.1	-0.3

^a ESX-: Essex Cryogenics 34 L stainless steel cylinder; GM-24/: Graeven Metalltechnik 27 L stainless steel cylinder; GM-26/Graeven Metalltechnik 50 L stainless steel cylinder; Lin-: Linde 50 L stainless steel cylinder; Lux-: Luxfer UK 50 L aluminium cylinder.

^b CSIRO94 assignments done by MPI-BGC. See text.

^c σ : standard error of the mean of the average analysis results of the individual measurement rounds in the period Sept-Dec2008.

^d Residuals of the calibration of the GC-PDD instrument: (H₂ assigned) – (measured H₂).

Table 2. H₂-in-air mixtures prepared by the mixing procedure.

SampleID	Production date	Analysis date	H ₂ [ppb]	signal ^a	σ^b [%]	n^c	residual ^d		u_{mix}^e [%]	
							[ppb]	[%]	random	total
20083923	26Nov08	27Nov08	12.4	0.021	10.7	18			0.3	0.39
20083894	25Nov08	25Nov08	49.8	0.068	1.7	5	0.1	0.21	0.19	0.32
20084077	09Dec08	09Dec08	99.2	0.141	1.2	14	1.2	1.2	0.26	0.36
20083352	16Oct08	16Oct08	142.9	0.214	1.2	20	-0.6	0.39	0.15	0.30
20084065	08Dec08	09Dec08	158.1	0.237	1.3	20	1.3	0.84	0.25	0.35
20083893	24Nov08	25Nov08	171.8	0.262	1.0	37	0.4	0.23	0.15	0.30
20083671	03Nov08	05Nov08	192.7	0.300	1.1	68	-0.1	0.07	0.17	0.31
20083557	27Oct08	28Oct08	223.2	0.356	0.87	28	0.1	0.04	0.18	0.31
20083261	07Oct08	08Oct08	265.9	0.441	0.67	14	-1.5	0.58	0.15	0.30
20084105	12Dec08	15Dec08	269.9	0.446	0.55	47	-0.1	0.02	0.16	0.30
20083408	17Oct08	17Oct08	287.8	0.485	0.65	17	-1.6	0.57	0.17	0.31
20083534	23Oct08	24Oct08	317.9	0.543	0.53	25	0.1	0.04	0.19	0.31
20083450	20Oct08	20Oct08	318.1	0.543	0.49	30	0.5	0.17	0.17	0.31
20083787	11Nov08	14Nov08	334.5	0.576	0.56	53	0.8	0.24	0.13	0.29
20083880	19Nov08	21Nov08	349.0	0.610	0.44	33	-0.5	0.14	0.18	0.31
20083282	09Oct08	09Oct08	361.2	0.637	0.49	38	-1.1	0.30	0.17	0.30
20083348	14Oct08	15Oct08	382.8	0.680	0.38	40	0.3	0.08	0.16	0.30
20083295	13Oct08	13Oct08	414.1	0.750	0.38	20	-0.1	0.03	0.13	0.28
20083670	30Oct08	03Nov08	438.8	0.806	0.34	59	-0.5	0.12	0.14	0.29
20083554	27Oct08	27Oct08	452.6	0.832	0.44	25	1.4	0.32	0.14	0.29
20083821	17Nov08	18Nov08	472.1	0.879	0.36	27	0.0	0.00	0.18	0.31
20083351	15Oct08	16Oct08	472.1	0.882	0.42	39	-1.0	0.21	0.08	0.27

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Table 2. Continued.

SampleID	Production date	Analysis date	H ₂ [ppb]	signal ^a	σ^b [%]	n^c	residual ^d		u_{mix}^e [%]	
							[ppb]	[%]	random	total
20083533	22Oct08	23Oct08	498.0	0.935	0.40	34	1.5	0.29	0.15	0.29
20083260	07Oct08	08Oct08	527.0	1.004	0.40	39	0.5	0.09	0.12	0.28
20084012	08Dec08	08Dec08	540.3	1.037	0.27	11	-0.5	0.09	0.19	0.31
20083114	25Sep08	26Sep08	540.4	1.037	0.38	32	-0.2	0.04	0.08	0.27
20083733	06Nov08	09Nov08	555.3	1.071	0.30	61	0.2	0.03	0.19	0.32
20083449	17Oct08	20Oct08	576.7	1.121	0.45	44	0.1	0.01	0.08	0.27
20083462	21Oct08	21Oct08	598.1	1.171	0.32	27	0.1	0.01	0.08	0.27
20083891	21Nov08	23Nov08	626.2	1.237	0.33	40	0.2	0.04	0.08	0.27
20083115	26Sep08	29Sep08	627.2	1.242	0.47	52	-1.0	0.16	0.08	0.27
20083786	11Nov08	11Nov08	647.7	1.286	0.20	34	1.1	0.17	0.08	0.27
20083028	16Sep08	17Sep08	668.8	1.340	0.35	39	-0.5	0.07	0.08	0.27
20083062	17Sep08	19Sep08	670.7	1.346	0.37	59	-1.2	0.18	0.08	0.27
20083553	24Oct08	27Oct08	695.5	1.400	0.31	68	1.2	0.17	0.08	0.27
20083820	14Nov08	17Nov08	724.6	1.472	0.34	58	0.5	0.07	0.14	0.29
20083668	28Oct08	29Oct08	751.7	1.537	0.31	43	0.3	0.05	0.08	0.27
20083860	17Nov08	20Nov08	778.9	1.600	0.28	39	1.7	0.21	0.14	0.29
20083711	06Nov08	07Nov08	805.6	1.667	0.19	34	1.1	0.14	0.15	0.29
20083954	28Nov08	01Dec08	825.0	1.716	0.24	51	0.3	0.04	0.15	0.29
20083297	14Oct08	14Oct08	849.4	1.779	0.35	19	-1.2	0.14	0.08	0.27
20083283	10Oct08	13Oct08	851.2	1.782	0.37	46	-0.5	0.05	0.08	0.27
20083669	29Oct08	30Oct08	868.1	1.828	0.29	45	-2.3	0.27	0.08	0.27
20082905	10Sep08	10Sep08	889.8	1.877	0.57	15	-0.6	0.07	0.08	0.27

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Table 2. Continued.

SampleID	Production date	Analysis date	H ₂ [ppb]	signal ^a	σ^b [%]	n ^c	residual ^d		u _{mix} ^e [%]	
							[ppb]	[%]	random	total
20083404	16Oct08	17Oct08	896.6	1.890	0.30	40	0.8	0.09	0.08	0.27
20084102	10Dec08	10Dec08	939.3	1.995	0.21	15	1.2	0.13	0.13	0.28
20084104	11Dec08	12Dec08	969.0	2.074	0.22	28	-1.2	0.12	0.11	0.28
20083710	05Nov08	06Nov08	1011	2.180	0.24	43	-1.5	0.15	0.09	0.27
20083816	14Nov08	14Nov08	1052	2.276	0.19	20	0.8	0.08	0.09	0.27
20083953	27Nov08	28Nov08	1099	2.399	0.36	46	-0.5	0.04	0.09	0.27
20084103	11Dec08	11Dec08	1181	2.602	0.18	19	1.4	0.12	0.09	0.27
20083850	18Nov08	19Nov08	1249	2.778	0.35	40	0.3	0.02	0.09	0.27
20084080	10Dec08	10Dec08	1252	2.787	0.31	24	-0.6	0.05	0.09	0.27

^a signal = average normalized peak height relative to the working standard containing H₂ = 523.8 ppb (SampleID 20061951),

^b σ = relative standard deviation of the analysis,

^c n = number of GC-HgO-analysis,

^d residual = H₂ mixing ratio measured based on the calibration function Eq. (3) – H₂ mixing ratio resulting from the moles of H₂ and moles of air mixed,

^e u_{mix} = uncertainty of the mixing procedure: random = combined uncertainty of pressure, temperature and mass measurements, total = combined uncertainty of all influencing parameters identified in Table 3.

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Table 3. Contributing factors to the uncertainty of the H₂ mixing procedure.

Variable	Source of error	uncertainty	relative uncertainty
p	sensor accuracy	$\pm 7 \text{ Pa}^a$	$0.01\%^1$
T	sensor accuracy	$\pm 0.2 \text{ K}^b$	$0.07\%^1$
	T inhomogeneity	$< \pm 0.1 \text{ K}^c$	$0.03\%^1$
	total T uncertainty	d	$0.08\%^1$
m (air)	balance accuracy	$\pm 0.16 \text{ g}^a$	$0.18\%^1$
ρ (air)	composition uncertainty	e	$< 0.01\%^2$
V	balance accuracy	$\pm 0.05 \text{ mg}^a$	$0.01\%^2$
	filling error	$\pm 0.3 \mu\text{L}^f$	$0.04\%^2$
	water density	$\pm 0.12 \text{ g/L}^g$	$0.01\%^2$
	internal valve vol.	$< \pm 0.5 \mu\text{L}^h$	$0.14\%^2$
	total V uncertainty	d	$0.15\%^2$
virial coefficient		$< 20\%^i$	$< 0.02\%^2$
Gas purity	H ₂ purity	$< 0.001\%^j$	$< 0.001\%^3$
	H ₂ blank of diluent	$< +1 \text{ ppb}^k$	$0.20\%^4$
	H ₂ from Al surface	$< +0.2 \text{ ppb}^l$	$< 0.04\%^4$
total error standard mixing		d	0.3%
regression fit error		0.5 ppb^m	0.1%
GC-HgO analysis		1.6 ppb^n	0.3%
total error at atmospheric mixing ratios		d	0.5%

^a DKD (German calibration service) calibration certificate; ^b checked with calibrated thermometer; ^c estimate from observed lab temperature variability; ^d root of sum of error squares; ^e estimate from literature (Park et al., 2004; Tohjima et al., 2005; NOAA, 1976); ^f reproducibility (1σ); ^g water density range of observed weighing lab temperature variability; ^h estimate of maximum additional fitting volume; ⁱ estimate based on the number of significant digits quoted in the literature; ^j manufacturer specification; ^k blank analysis using GC-HePDD; ^l drift over 5 days taken from storage experiment (s. Fig. 5); ^m see Table 1, last column; ⁿ scatter of daily averages of the quality control standard analysis (1σ) ¹ random and systematic influence; ² systematic bias; ³ bias resulting in systematically lower mixing ratios; ⁴ bias resulting in systematically higher mixing ratios.

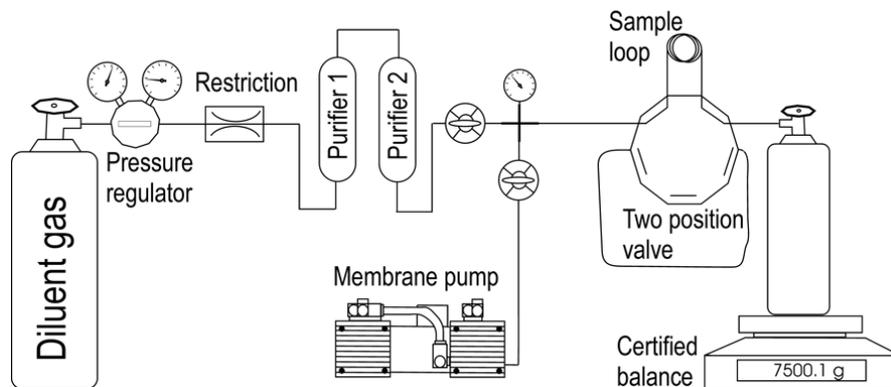


Fig. 1. Set-up of the system for mixing hydrogen and purified air.

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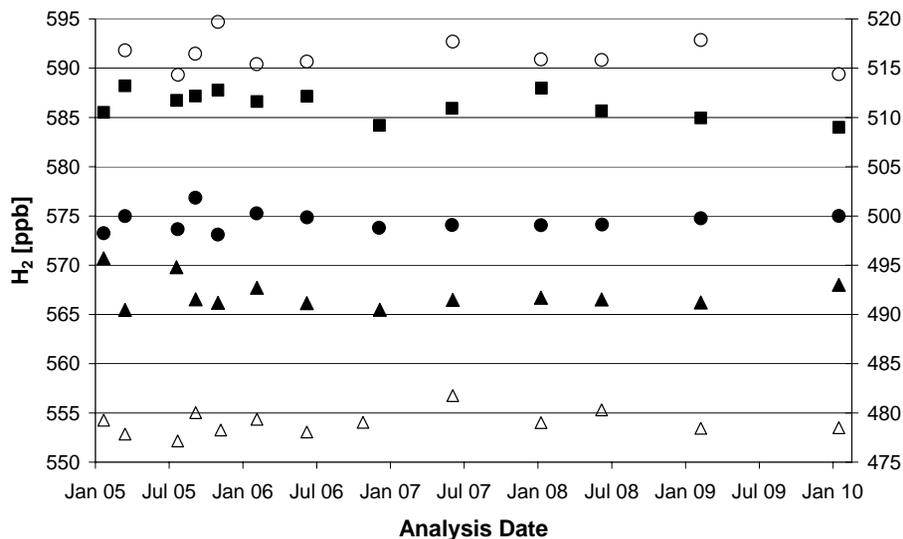


Fig. 2. Storage test of H₂ in air samples stored in glass flasks over five years. Filled symbols refer to the left axis, open symbols to the right axis.

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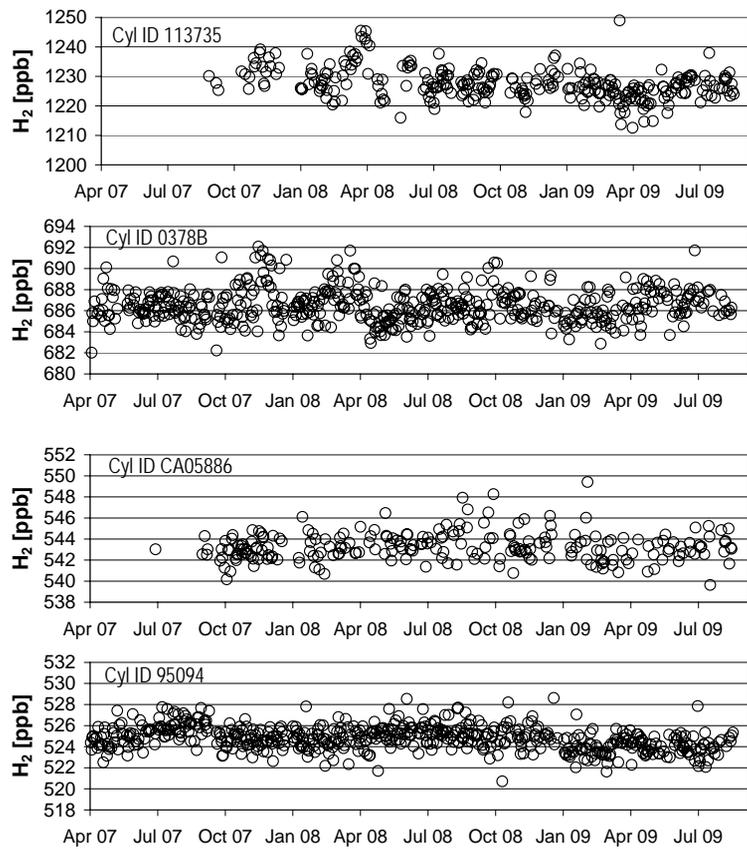


Fig. 3. Time series of quality control standard gas analysis results.

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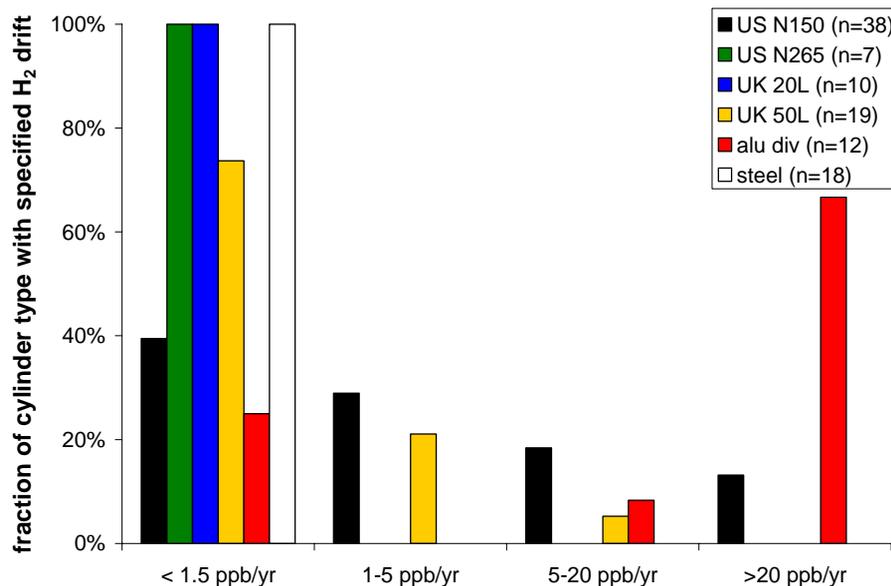


Fig. 4. Hydrogen drift rates in air stored in different high pressure cylinder types. The first two bars represent standards stored in Luxfer US aluminium cylinders (US N150 purchased from Scott Marrin Inc., Riverside (CA), USA; US N265 purchased from Conwin Carbonic Co., Los Angeles (CA), USA); blue and orange bars (UK 20L and 50 L) represent Luxfer UK aluminium cylinders (purchased from MATAR, Mazzano, Italy); further aluminium cylinders of various origins and sizes are grouped in the red column “alu div”.

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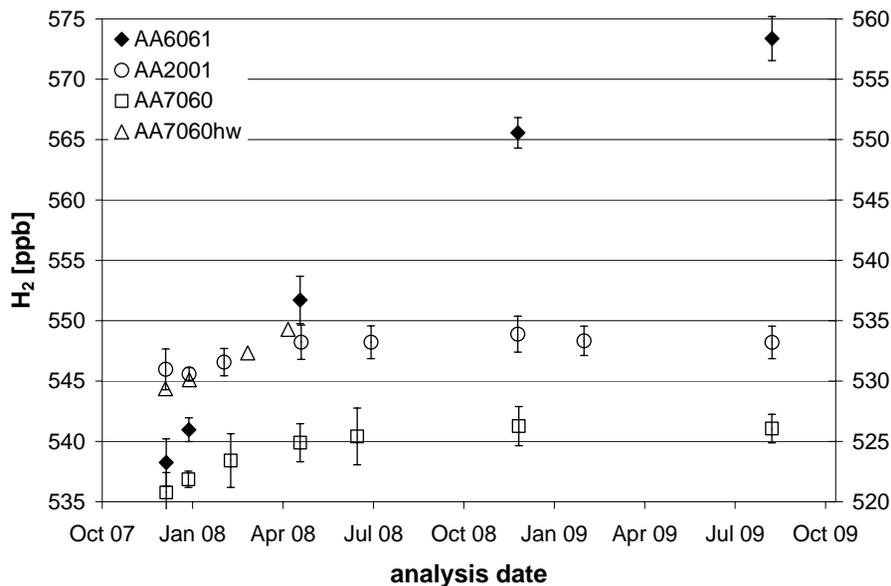


Fig. 5. Dependence of H₂ growth rate in air on aluminium cylinder alloy. Black circles refer to the left axis, open symbols refer to the right axis. Data of the cylinder used for the standard mixing procedure (AA7060hw) are illustrated by triangles.

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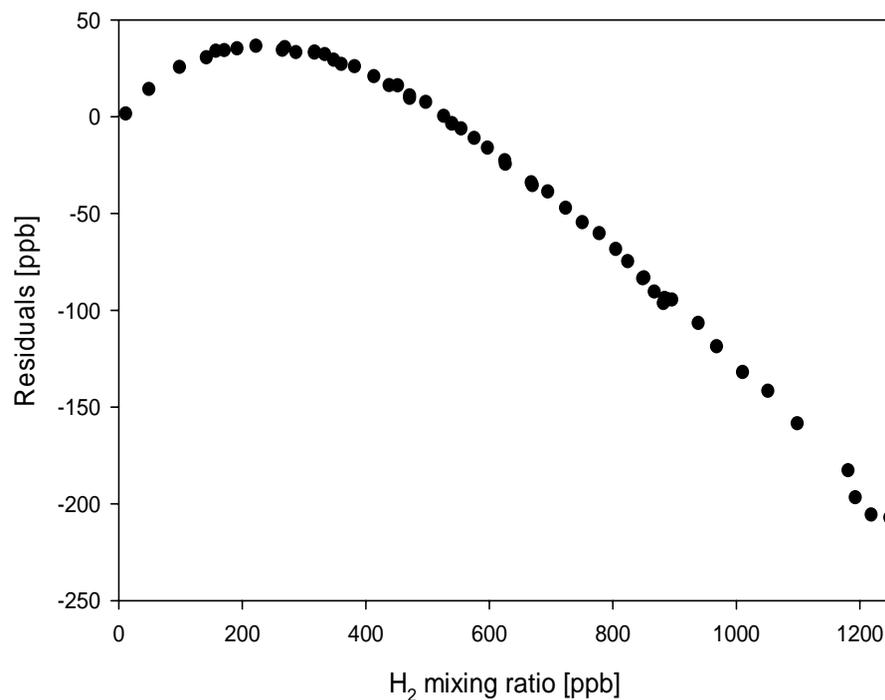
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Fig. 6. Residuals of a linear regression fit of the GC-HgO detector signals for the set of 53 H₂-in-air mixtures. Displayed is the difference: (H₂ amount mixed) – (measured H₂ mixing ratio) based on a one point calibration.

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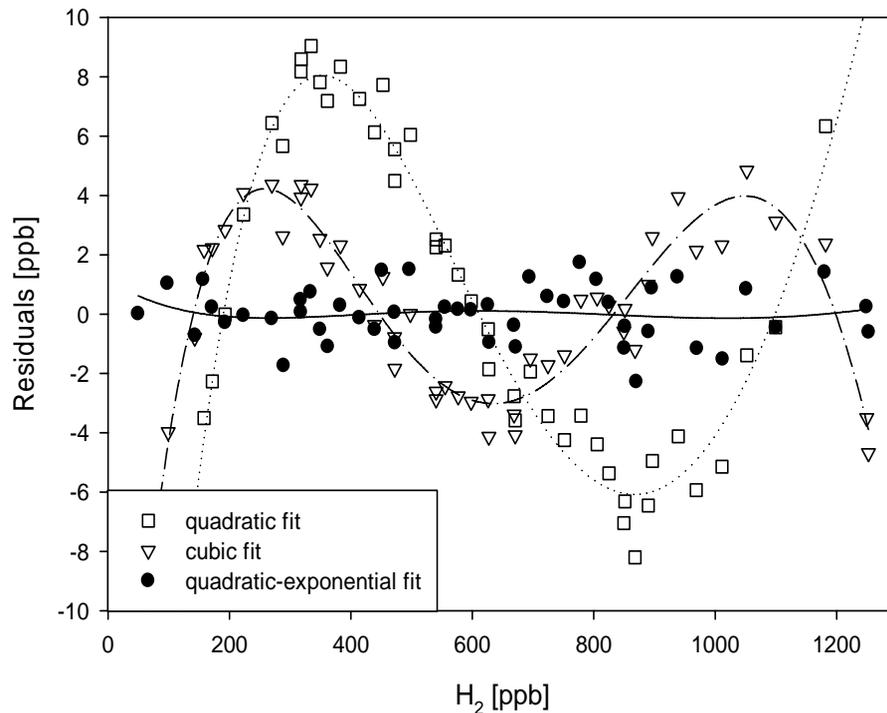
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Fig. 7. Residuals for various regression functions through the set of H_2 -in-air mixtures. Displayed is the difference: (H_2 amount mixed) – (measured H_2 mixing ratio) based on the respective regression function.

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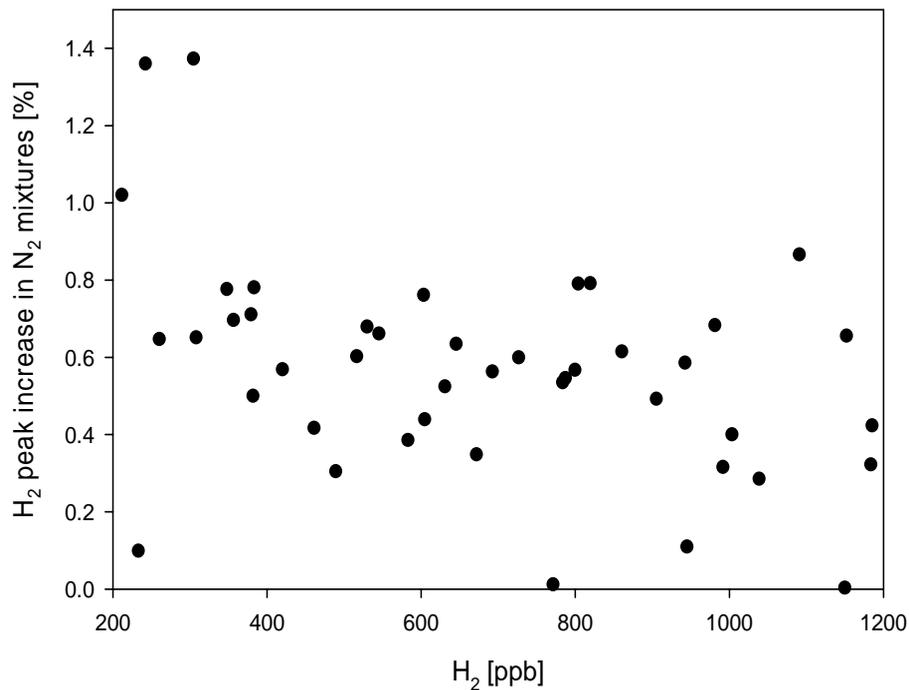
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Fig. 8. Relative increase of the normalized H₂ peak height in N₂ mixtures compared to H₂ in air mixtures [%].

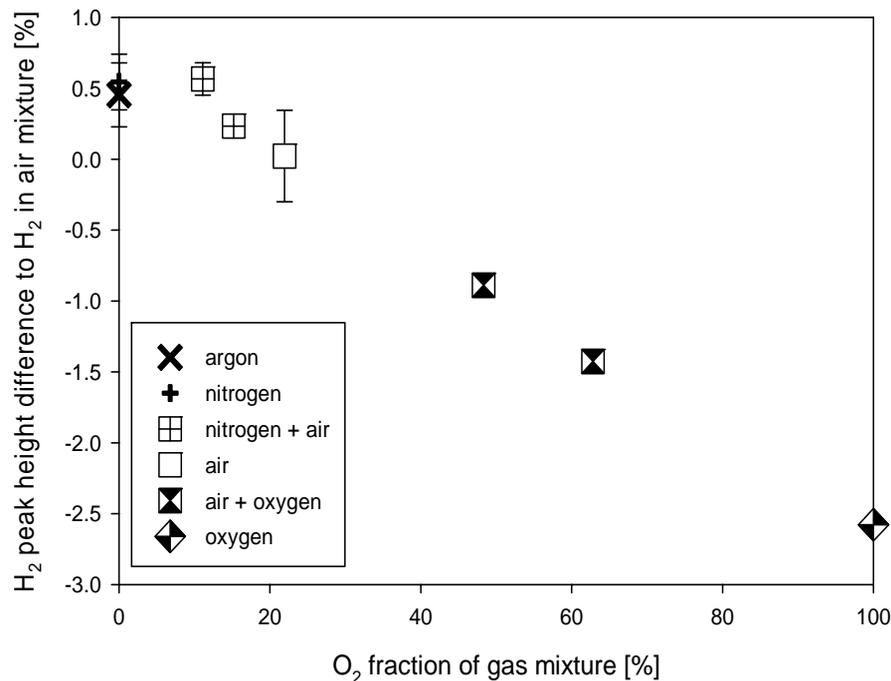
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Fig. 9. Dependence of the detector response on the oxygen content. Displayed are the differences between the assigned H₂ concentration based on the calibration with H₂-in-air mixtures and the H₂ concentration as it was mixed (extended from Jordan, 2009).

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